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## Colligative

### Properties Of

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Title: Colligative Properties of Solutions  
1 Colligative Properties of Solutions. Boiling Pt Elevation and Freezing Pt Depression; 2 Colligative Properties. depend only on the number of solute particles in a solution ; does not depend on the identity of particles; 3 Boiling Point Elevation. a nonvolatile solute lowers the vapor pressure

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## **PPT - Colligative Properties of Solutions**

**PowerPoint ...**

Colligative Properties -  
Solute particles take up  
space in a solution. ...

(density  $0.88 \text{ g cm}^{-3}$ )

in the osmometer

corrected for capillary

rise is  $11.6 \text{ cm}$  at ... |

PowerPoint PPT

presentation | free to

view. Colligative

Properties -  $P_A = .848$  (

$760 \text{ torr}$ )  $P_A = 644 \text{ torr}$

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...  
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**PPT - Colligative  
Properties  
PowerPoint  
presentation ...**

Colligative Properties  
Colligative Property: A  
property that depends  
only upon the number  
of solute particles  
(concentration), and  
not upon their identity.  
Three Important  
Colligative Properties  
of Solutions. Vapor-  
pressure lowering

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## **Chapter 16**

### **Colligative**

### **Properties of**

### **Solutions**

**|authorSTREAM**

1. Solutions Colligative Properties • Changes in colligative properties depend only on the number of solute particles present, not on the identity of the solute particles. •



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Among colligative properties are Vapor pressure lowering Boiling point elevation Melting point depression Osmotic Pressure. 2.

## **Colligative properties - LinkedIn SlideShare**

Colligative properties are those that depend on the concentration of particles in a solution, not upon the identity of those particles.

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**Solutions -**

**Presentation**

**Chemistry**

ppt on colligative properties of solutions are a good way to achieve details about operating certain products. Many products that you buy can be obtained using instruction manuals.

These user guides are clearly built to give step-

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by-step information about how you ought to go ahead in operating certain equipments.

## **PPT ON COLLIGATIVE PROPERTIES OF SOLUTIONS PDF**

- By definition a colligative property is a solution property (a property of mixtures) for which it is the amount of solute dissolved in the solvent matters but the kind of

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Properties Of

solute does not matter.

- Coming to grips with this concept should immediately remind you of kinetic molecular theory of gases—in that case we

**Colligative**

**Properties- Page 1**

**Lecture 4:**

**Colligative ...**

COLLIGATIVE

PROPERTIES Elevation  
of Boiling Point

Depression of Freezing

Point Lowering of

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Vapor Pressure  
Osmotic Pressure  
MOLE FRACTION &  
MOLALITY MOLE  
FRACTION OF

Component  $i = X_i = n_i$   
/  $n_{\text{total}}$  (c.f Gases;  
Chapter 5, p.217)

MOLALITY = Moles of  
Solute / kg Solvent

MOLALITY Useful when  
Temperature Changes  
are considered, as  
Volumes of solutions  
change with changing  
temperature, whereas

...

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**COLLIGATIVE  
PROPERTIES**

Among colligative  
properties are Vapor  
pressure lowering  
Boiling point elevation  
Melting point  
depression Osmotic  
pressure © 2009,  
Prentice-Hall, Inc.

Vapor Pressure  
Because of solute-  
solvent intermolecular  
attraction, higher  
concentrations of  
nonvolatile solutes

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make it harder for

solvent to escape to

the vapor phase. ©

2009, Prentice-Hall,

Inc. Raoult's Law  $P_A =$

$X_A P_A^*$  where  $X_A$  is the

mole fraction of

compound A, and  $P_A^*$

is the normal vapor

pressure of A at that

temperature.

## **Chapter 13**

### **Properties of**

### **Solutions - Colby**

### **College**

Properties of solutions

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that depend on the number of molecules present and not on the kind of molecules are called colligative properties. These properties include boiling point elevation, freezing point depression, and osmotic pressure.

**Colligative  
Properties -  
University Of  
Cincinnati**

Colligative Properties



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## Colligative

### Properties Of

Of Solutions. 1.

Colligative Properties  
of Solutions<br />. 2.

The presence of solutes affects the properties of solutions.<br />Some of these properties are not dependant on what dissolved substance but only on how much.<br />Colligative properties are those that depend on the concentration of solute particles but not on their identity.<br

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Colligative Properties  
of Ionic Solutions

Colligative properties  
of solutions depends  
upon the total  
concentration of  
particles. Each  
equation describing  
colligative properties  
must be modified to

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account for this with ionic solutions since each ionic compound gives more than one mole of ions for every mole of compound.

## **PowerPoint Presentation**

Osmotic pressure is the most important of the colligative properties since it is related to the physiological compatibility of parenteral, ophthalmic and nasal solutions.

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Properties Of

**A Lecture on**

**Colligative**

**Properties in an**

**Undergraduate ...**

Colligative properties of solutions are properties that depend upon the concentration of solute molecules or ions, but not upon the identity of the solute. Colligative properties include vapor pressure lowering, boiling point elevation, freezing point depression, and

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osmotic pressure.

Lowering the Vapor

Pressure:

## **Colligative Properties - Chemistry & Biochemistry**

Colligative Properties

Properties of solutions

(compared to the pure

solvent) Depend upon

the number of solute

molecules or ions

(concentration), but

NOT upon the identity

of the solute. Include

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freezing point depression, boiling point elevation, vapor pressure, and osmotic pressure. You should be able to: 1. Define the CP 2.

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Colligative Properties

A. Definition Colligative

Property • property

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## Colligative

### Properties Of

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that depends on the

number (amount) of

solute particles in the

solution & not in the

nature of the solute

particles Colligative

Properties Vapor

Pressure Lowering

Boiling Point Elevation

Freezing Point

Depression Osmotic

Pressure Vapor

Pressure Lowering

Vapor Pressure-

pressure exerted by

the vapor/gas above

the liquid VP of sol'n <



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Properties Of

VP of solvent  $P = P_1^0 -$   
 $P_1 = X_2 P_1^0$  where:  $P_1^0$   
 $=$  V.P. of solvent  $X_2 =$   
mole frxn of ...

**colligative.ppt -**  
**Solutions I II III III**  
**Colligative ...**

Colligative Properties A dilute solution is one in which the amount of the solute is very small in comparison to the amount of the solvent. The dilute solutions show more or less ideal behavior as

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the heat and volume changes, accompanying the mixing of solute and solvent, are negligible for all practical purposes.

Copyright code: d41d8  
cd98f00b204e9800998  
ecf8427e.