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atom is a tiny,  
indestructible particle

with no internal  
structure. Thomson.

The atom is a sphere of  
positive electrical

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CHAPTER 5 What You'll

Learn You will compare

the wave and particle

models of light. You

will describe how the

frequency of light

emitted by an atom is

a unique characteristic

of that atom. You will

compare and contrast

the Bohr and quantum

mechanical models of

the atom. You will

express the

arrangements of

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electrons in atoms  
through orbital

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Jennie L. Borders.

Section 5.1 - Models of  
the Atom. The

Rutherford's model of  
the atom did not  
explain how an atom  
can emit light or the



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chemical properties of an atom. Plum Pudding Model Rutherford's Model. The Bohr Model.

### Answers

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...are the way electrons are arranged in various orbitals around the nuclei of atoms. Three rules tell us how: Aufbau principle - electrons enter the lowest energy first. This causes difficulties because of the overlap

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of orbitals of different energies - follow the diagram! Pauli Exclusion Principle - at most 2 electrons per orbital ...

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Chapter 5 - electrons in atoms (handouts)

Chapter 6 - periodic table & trends

(handouts) Chapters 7/9 - ionic bonding & naming (handouts)

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Chapters 8/9 - covalent bonding & chemical names & formulas (handouts) Chapters

8/15 - VSEPR/polar bonding/IMFs (handouts) Chapter 10

- moles (handouts) Chapter 11 - reactions (handouts) Chapter 12

...

...

## **Science / Chapter 5 - electrons in atoms (handouts)**

Chapter 5: Electrons in Atoms Models of the

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Atom Rutherford used existing ideas about the atom and proposed an atomic model in which the electrons move around the nucleus, like the planets move around the sun. Rutherford's model fails to explain why objects change color when heated.

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Terms in this set (82)

Rutherford's model-

also called nuclear

model-did not explain

how the electrons were

arranged

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the Atom. The

Rutherford's model of  
the atom did not

explain how an atom

can emit light or the

chemical properties of

an atom. Plum Pudding

Model Rutherford's

Model. The Bohr Model.

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Electrons in Atoms

Section 55.1.1 Figure

5.1 Different elements  
can have similar

reactions with water.

Objectives Compare

the wave and particle

natures of light. Define

a quantum of energy,

and explain how it is

related to an energy



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Atoms  
change of matter.  
Contrast continuous  
electromagnetic

Assessment

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IB Chapter Questions.

02:25. Problem 1

Objects get their colors  
from reflecting only  
certain wavelengths  
when hit with white  
light. Light reflected  
from a green leaf is  
found to have a

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wavelength of  $4.90$

$\times 10^{-7}$

$\text{m}$ . What

is the frequency of the

light? ...

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How are electrons  
arranged in an atom?

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times. 0. 10th grade .

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5 Electrons in Atoms  
Quiz - Quizizz**

Chapter 5“Electrons in  
Atoms ... Describe the  
energies and positions  
of electrons according  
to the quantum  
mechanical model.

Section 5.1 Models of

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### Electrons In

the Atom. OBJECTIVES:

Describe how the shapes of orbitals related to different sublevels differ. Ernest Rutherford's Model.

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How many electrons can each p orbital

hold? Chapter 5:

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